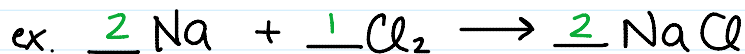


6.1 Types of Chemical Reactions

Tuesday, March 28, 2017
11:40 AM

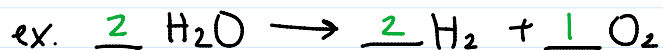
① Synthesis (S) Reaction

→ 2 or more reactants/elements combine to produce a single product



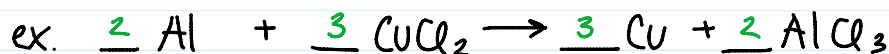
② Decomposition (D) Reaction

→ the breaking down of a compound into separate elements.



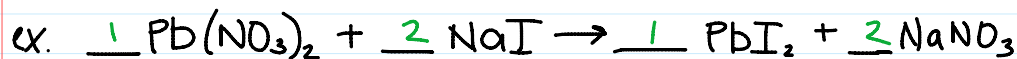
③ Single Replacement (SR) Reaction

→ replace 1 element from a compound with a separate element added as a reactant.



④ Double Replacement (DR) Reaction

→ swap elements between 2 compounds
→ they react together to form 2 new compounds



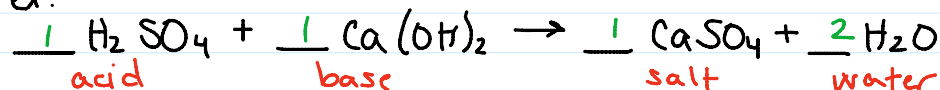
⑤ Neutralization (N) or Acid-Base Reactions

→ an acid and base react to form salt and water

ACID: most compounds starting w/ H (hydrogen)

BASE: most compounds ending in OH or beginning with NH₄

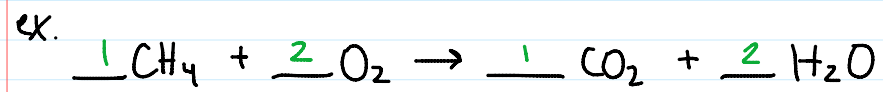
ex.



⑥ Combustion (C) Reactions

→ compound/element reacts with oxygen to form an oxide and heat energy

ex.



ex.

