**C:\Documents and Settings\gill_narinder\Local Settings\Temporary Internet Files\Content.IE5\FIBIMOIY\MC900250132[1].wmf**Naming Multivalent Compounds

Name the following compounds. DON’T forget the Roman Numerals!

|  |  |  |  |
| --- | --- | --- | --- |
| **FORMULA** | **ROUGH WORK** | **NAME** | **TOTAL # OF ATOMS** |
| 1. Hg2S |  |  |  |
| 1. CuI2 |  |  |  |
| 1. PbS |  |  |  |
| 1. PbCl2 |  |  |  |
| 1. PbCl4 |  |  |  |
| 1. CoCl2 |  |  |  |
| 1. AuCl |  |  |  |
| 1. AuCl3 |  |  |  |
| 1. HgO |  |  |  |
| 1. CuCl2 |  |  |  |
| 1. Fe2O3 |  |  |  |
| 1. Cu2O |  |  |  |

**C:\Documents and Settings\gill_narinder\Local Settings\Temporary Internet Files\Content.IE5\FIBIMOIY\MC900250132[1].wmf**Writing Multivalent Compounds

Write the chemical formulae for the following compounds. Show all your work!

|  |  |  |
| --- | --- | --- |
| **CHEMICAL NAME** | **FORMULA** | **TOTAL # OF ATOMS** |
| 1. Tin (II) oxide |  |  |
| 1. Lead (IV) fluoride |  |  |
| 1. Iron (III) sulphide |  |  |
| 1. Chromium (II) iodide |  |  |
| 1. Lead (IV) bromide |  |  |
| 1. Nickel (III) chloride |  |  |
| 1. Mercury (II) chloride |  |  |
| 1. Gold (III) nitride |  |  |
| 1. Tin (IV) oxide |  |  |
| 1. Lead (IV) sulphide |  |  |
| 1. Titanium (III) oxide |  |  |
| 1. Vanadium (VI) phosphide |  |  |
| 1. Chromium (II) oxide |  |  |
| 1. Manganese (II) oxide |  |  |
| 1. Cobalt (III) bromide |  |  |
| 1. Copper (II) nitride |  |  |
| 1. Palladium (IV) oxide |  |  |
| 1. Palladium (II) oxide |  |  |
| 1. Copper (I) nitride |  |  |
| 1. Cobalt (II) chloride |  |  |