

Standard Atomic Notation

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11:37 AM

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Quick review:

| | P^+ | N^0 | E^- |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----------------------|------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------|
| <div style="border: 1px solid black; padding: 5px; display: inline-block;">16 2^- S 32.1</div> | = atomic # = 16 | = round off atomic mass to get mass # = 32 mass # - P^+ = $32 - 16$ = 16 | = atomic # - charge = $16 - (-2)$ = $16 + 2$ = 18 |

Isotopes: - atoms of the same element that have different mass #'s.

- the atomic mass of an element is the average mass of an elements naturally occurring atoms

- 2 common ways to refer to isotopes are with the **name** and **symbol**.

ex. Uranium

name: Uranium - 238

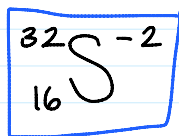
refers to the mass #

symbol: ^{238}U

atomic #

ex. sulfur ion

symbol:



ex. Chlorine atom

symbol:

